**Revision stuff for Test #4 IMF/ solutions test**

1. Revise the order of increasing polarity of bond. Revise NC page: 264 and 265

Look at F2, Cl F, BrF, IF, Co2, So2, CCl4

1. Order of increasing melting point in graphite, O2, C3H8 and ethanol
2. Revise 268- 272
3. Page: 280 graph, solubility curves and worked example: 6.1
4. Saturated, unsaturated and supersaturated solutions
5. Concentration, moles, mass calculations
6. Rf value of most polar component in TLC
7. Draw the structure, shape, name of shape and polarity of molecule. Learn page: 262
8. Compare the vapour pressure and density of liquid water to the other named substances. Justify the reasoning behind your prediction, using diagrams where appropriate.
9. Compared to pentane (C5H12), the **vapour pressure** of liquid water is: Compared to solid water (ice), the **density** of liquid water is(reason & diagram)
10. sodium chloride dissolve in polar solvents like formaldehyde. Name the type of attractive force that allows sodium chloride to dissolve in formaldehyde.
11. Which has highest boiling point? SO2or CO2Reason. Bromine/ iodine
12. Explain why some substances are described as polar molecules. What force is present when NaCl dissolves in a polar solvent?
13. Explain why water has high boiling point compared to methane?

**In addition complete chapter questions and worksheets.**

